H = enthalpy
P = pressure
R = universal gas constant, 1.9872 B.t.u./lb.-mole ° R.
S = entropy
T = temperature
V = volume
α = residual volume

Superscripts

° = zero pressure, ideal gas state
* = unit fugacity, ideal gas state

Subscripts

c = critical constants
L = liquid phase
v = vapor phase

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CORRECTION

In "Prediction of Equilibrium Ratios from Nomograms of Improved Accuracy" [B.C. Cajander, H.G. Hipkin, and J.M. Lenoir, J. CHEM. ENG. DATA **5**, 251 (1960)] the figures entitled "Equilibrium ratios of aliphatic hydrocarbons at 10 p.s.i.a." in the right column on page 254 are incorrect. The line labeled butadiene-1,2 is low by about 30%. The line labeled 2 or 3-methylbutene-1 is correct for 3-methylbutene-1, but 2-methylbutene-1 should lie between pentene-1 and isoprene. The corrected figures appear as shown below.